

# Cambridge International AS & A Level

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**CHEMISTRY****9701/12**

Paper 1 Multiple Choice

**February/March 2026****1 hour 15 minutes**

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet  
Soft clean eraser  
Soft pencil (type B or HB is recommended)

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**INSTRUCTIONS**

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A, B, C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid or tape.
- Do **not** write on any bar codes.
- You may use a calculator.

**INFORMATION**

- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.
- Important values, constants and standards are printed in the question paper.

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This document has **20** pages. Any blank pages are indicated.

1 Which statement about p orbitals is correct?

- A** All p orbitals are  shaped.
- B** All p orbitals have a lower energy than an s orbital of the same principal quantum number.
- C** Each of the first four quantum shells contain p orbitals.
- D** Each p orbital can contain a maximum of six electrons.

2 The first four ionisation energies for phosphorus, in  $\text{kJ mol}^{-1}$ , are shown.

| 1st  | 2nd  | 3rd  | 4th  |
|------|------|------|------|
| 1060 | 1900 | 2920 | 4960 |

Which row gives the 5th to 14th ionisation energies for phosphorus?

|          | 5th    | 6th    | 7th    | 8th    | 9th    | 10th   | 11th   | 12th   | 13th    | 14th    |
|----------|--------|--------|--------|--------|--------|--------|--------|--------|---------|---------|
| <b>A</b> | 6270   | 8410   | 21 270 | 25 400 | 29 800 | 35 870 | 40 100 | 46 300 | 59 000  | 271 800 |
| <b>B</b> | 6270   | 21 270 | 25 400 | 29 860 | 35 870 | 40 100 | 46 300 | 54 070 | 59 000  | 271 800 |
| <b>C</b> | 6270   | 21 270 | 25 400 | 29 860 | 35 870 | 40 100 | 46 300 | 59 000 | 271 800 | 301 400 |
| <b>D</b> | 21 270 | 25 400 | 29 860 | 35 870 | 40 100 | 46 300 | 54 070 | 59 000 | 271 800 | 301 400 |

3 Sodium carbonate neutralises phosphoric acid to produce sodium phosphate, carbon dioxide and water only.

Which volume of  $0.50 \text{ mol dm}^{-3}$  sodium carbonate is needed to exactly neutralise  $40.0 \text{ cm}^3$  of  $0.25 \text{ mol dm}^{-3}$  phosphoric acid?

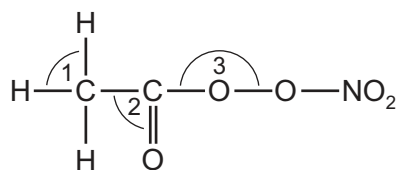
- A**  $10.0 \text{ cm}^3$       **B**  $13.3 \text{ cm}^3$       **C**  $20.0 \text{ cm}^3$       **D**  $30.0 \text{ cm}^3$

4 Which equation describes the complete combustion of an alkene,  $\text{C}_n\text{H}_{2n}$ ?

- A**  $\text{C}_n\text{H}_{2n} + 2n\text{O}_2 \rightarrow n\text{CO}_2 + n\text{H}_2\text{O}$
- B**  $\text{C}_n\text{H}_{2n} + 2n\text{O}_2 \rightarrow n\text{CO}_2 + 2n\text{H}_2\text{O}$
- C**  $\text{C}_n\text{H}_{2n} + 1.5n\text{O}_2 \rightarrow n\text{CO}_2 + 2n\text{H}_2\text{O}$
- D**  $\text{C}_n\text{H}_{2n} + 1.5n\text{O}_2 \rightarrow n\text{CO}_2 + n\text{H}_2\text{O}$

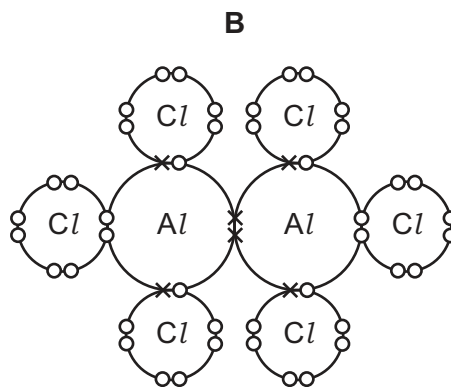
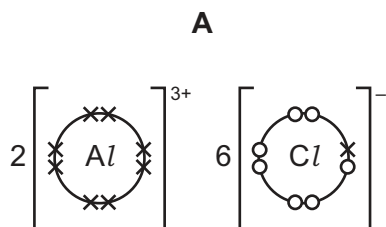
- 5 Organic nitrates in photochemical smog can cause breathing difficulties.

The diagram shows an example of an organic nitrate molecule.



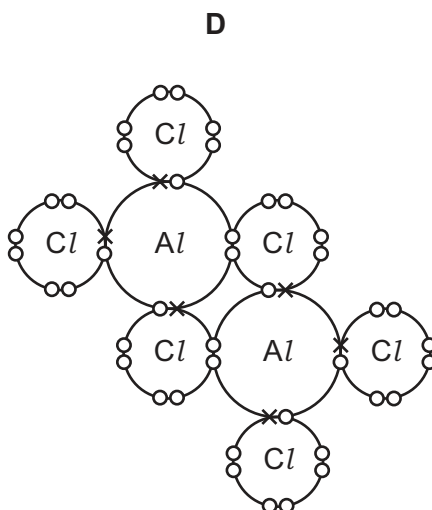
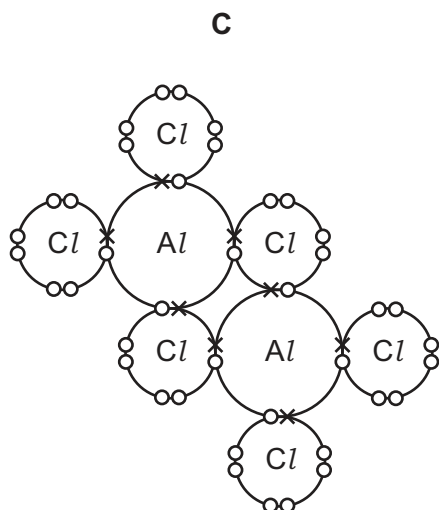
What is the correct order of the bond angles shown in ascending order (smallest first)?

- A  $1 \rightarrow 2 \rightarrow 3$       B  $2 \rightarrow 1 \rightarrow 3$       C  $3 \rightarrow 1 \rightarrow 2$       D  $3 \rightarrow 2 \rightarrow 1$
- 6 Which dot-and-cross diagram is correct for  $Al_2Cl_6$ ?



**key**

- × electrons from  $Al$   
 ○ electrons from  $Cl$



- 7 In this question you should assume methane behaves as an ideal gas.

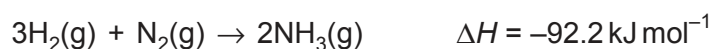
The gas laws can be summarised in the ideal gas equation.

$$pV = nRT$$

The volume of a sample of methane is measured at a temperature of 60 °C and a pressure of 103 kPa. The volume measured is  $5.37 \times 10^{-3} \text{ m}^3$ .

What is the mass of the sample of methane, given to two significant figures?

- A** 0.0032 g      **B** 0.018 g      **C** 3.2 g      **D** 18 g
- 8 The equation for the formation of ammonia is shown.

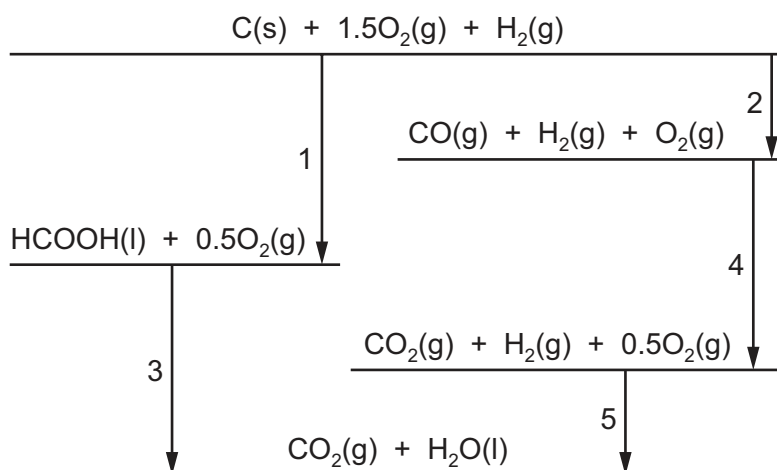


Some bond energy data is given.

| bond | bond energy / $\text{kJ mol}^{-1}$ |
|------|------------------------------------|
| H–H  | 436                                |
| N–H  | 388                                |

What is the bond energy of the  $\text{N}\equiv\text{N}$  bond in the nitrogen molecule?

- A** 248  $\text{kJ mol}^{-1}$       **B** 928  $\text{kJ mol}^{-1}$       **C** 1392  $\text{kJ mol}^{-1}$       **D** 1848  $\text{kJ mol}^{-1}$
- 9 The diagram shows an energy cycle with a series of reactions.



Which reactions have an enthalpy change that is an enthalpy of formation?

- A** 1, 2 and 5      **B** 1 only      **C** 2, 3 and 4      **D** 3, 4 and 5

10 Which reaction is an example of disproportionation?

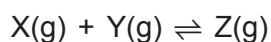
- A  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$   
 B  $2\text{KHCO}_3 \rightarrow \text{K}_2\text{CO}_3 + \text{CO}_2 + \text{H}_2\text{O}$   
 C  $2\text{KMnO}_4 \rightarrow \text{K}_2\text{MnO}_4 + \text{MnO}_2 + \text{O}_2$   
 D  $2\text{NaNO}_3 \rightarrow 2\text{NaNO}_2 + 3\text{O}_2$

11 Compound W contains chlorine in an oxidation state of +7.

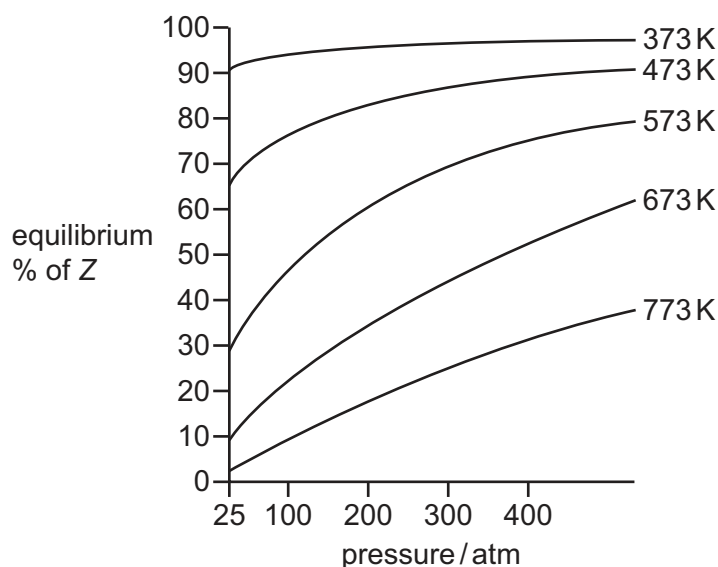
What is compound W?

- A  $\text{NH}_4\text{Cl}$       B  $\text{NH}_4\text{ClO}_3$       C  $\text{NH}_4\text{ClO}_4$       D  $\text{N}_2\text{H}_5\text{Cl}$

12 In an industrial process, two gases X and Y react together to form a single gaseous product Z.



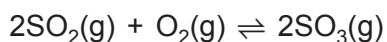
The equilibrium percentage of product Z varies according to the pressure and the temperature as shown in the graphs.



Which statement about this equilibrium reaction is correct?

- A Decreasing the temperature decreases the value of the equilibrium constant.  
 B Decreasing the temperature increases the rate of this reaction.  
 C Increasing the pressure increases the value of the equilibrium constant.  
 D The reaction is exothermic in the forward direction.

- 13 The Contact process involves an equilibrium reaction.



Which row gives the  $K_p$  expression and the pressure conditions used for this reaction?

|          | $K_p$ expression                                           | pressure / atm |
|----------|------------------------------------------------------------|----------------|
| <b>A</b> | $\frac{p^2\text{SO}_3}{p^2\text{SO}_2 \times p\text{O}_2}$ | 2              |
| <b>B</b> | $\frac{p^2\text{SO}_2 \times p\text{O}_2}{p^2\text{SO}_3}$ | 2              |
| <b>C</b> | $\frac{p^2\text{SO}_3}{p^2\text{SO}_2 \times p\text{O}_2}$ | 200            |
| <b>D</b> | $\frac{p^2\text{SO}_2 \times p\text{O}_2}{p^2\text{SO}_3}$ | 200            |

- 14 When nitrogen monoxide gas is mixed with oxygen gas, the molecules collide with each other. Some of these collisions are effective and lead to a reaction.

The pressure of the mixture was increased at constant temperature.

Which row correctly describes how the frequency of effective and non-effective collisions changed?

|          | frequency of effective collisions | frequency of non-effective collisions |
|----------|-----------------------------------|---------------------------------------|
| <b>A</b> | increased                         | decreased                             |
| <b>B</b> | increased                         | increased                             |
| <b>C</b> | increased                         | unchanged                             |
| <b>D</b> | unchanged                         | decreased                             |

- 15 A reversible chemical reaction is studied.

The forward reaction has an activation energy  $E_a = 50 \text{ kJ mol}^{-1}$  and an enthalpy change of reaction  $\Delta H_f = +20 \text{ kJ mol}^{-1}$ .

When a catalyst is added, the activation energy of the forward reaction decreases by 50%.

What is the activation energy of the backwards reaction when the catalyst is present?

- A**  $45 \text{ kJ mol}^{-1}$       **B**  $25 \text{ kJ mol}^{-1}$       **C**  $15 \text{ kJ mol}^{-1}$       **D**  $5 \text{ kJ mol}^{-1}$

16 Which substance does **not** give  $\text{HCl(g)}$  when water is added to it?

- A**  $\text{MgCl}_2$                       **B**  $\text{Al}_2\text{Cl}_6$                       **C**  $\text{SiCl}_4$                       **D**  $\text{PCl}_5$

17 For Groups 1 and 2, metallic bonding is stronger in metals that have a smaller atomic radius.

X and Y are both metals.

$\Delta H_{\text{at}}$  is the energy required to produce 1.0 mol of gaseous atoms from the solid.

$$\Delta H_{\text{at}}\text{X} = +107 \text{ kJ mol}^{-1}; \Delta H_{\text{at}}\text{Y} = +148 \text{ kJ mol}^{-1}$$

What are X and Y?

|          | <b>X</b> | <b>Y</b> |
|----------|----------|----------|
| <b>A</b> | Na       | K        |
| <b>B</b> | Na       | Mg       |
| <b>C</b> | Mg       | Na       |
| <b>D</b> | Mg       | K        |

18 The elements in Period 3 and their compounds show trends across the period from sodium to sulfur.

Which row is correct?

|          | <b>electronegativity of the elements</b> | <b>acid–base behaviour of the oxides of the elements</b> |
|----------|------------------------------------------|----------------------------------------------------------|
| <b>A</b> | decreases                                | basic → amphoteric → acidic                              |
| <b>B</b> | decreases                                | acidic → amphoteric → basic                              |
| <b>C</b> | increases                                | basic → amphoteric → acidic                              |
| <b>D</b> | increases                                | acidic → amphoteric → basic                              |

- 19 This question is about the hydroxides and carbonates of magnesium and barium.

Which row is correct?

|   | the more soluble hydroxide | the carbonate with the greater thermal stability |
|---|----------------------------|--------------------------------------------------|
| A | Ba(OH) <sub>2</sub>        | BaCO <sub>3</sub>                                |
| B | Mg(OH) <sub>2</sub>        | BaCO <sub>3</sub>                                |
| C | Ba(OH) <sub>2</sub>        | MgCO <sub>3</sub>                                |
| D | Mg(OH) <sub>2</sub>        | MgCO <sub>3</sub>                                |

- 20 When 3.00 g of an anhydrous nitrate of a Group 2 metal is decomposed, 1.53 g of gas is produced.

What is the nitrate compound?

- A beryllium nitrate  
 B calcium nitrate  
 C magnesium nitrate  
 D strontium nitrate
- 21 When concentrated sulfuric acid is added to crystals of sodium bromide, the products include a colourless gas and a brown-coloured substance.

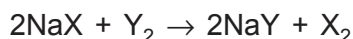
The colourless gas gives a cream precipitate with AgNO<sub>3</sub>(aq).

Which row describes the reactions that produce the colourless gas and the brown-coloured substance?

|   | reaction that produces colourless gas | reaction that produces brown-coloured substance |
|---|---------------------------------------|-------------------------------------------------|
| A | acid–base                             | acid-base                                       |
| B | acid-base                             | redox                                           |
| C | redox                                 | acid-base                                       |
| D | redox                                 | redox                                           |

22 X and Y are halogens.

The following reaction occurs in an aqueous solution.



Which row is correct?

|   | X        | Y        | behaviour of NaX      |
|---|----------|----------|-----------------------|
| A | bromine  | chlorine | as an oxidising agent |
| B | bromine  | chlorine | as a reducing agent   |
| C | chlorine | bromine  | as an oxidising agent |
| D | chlorine | bromine  | as a reducing agent   |

23 Mohr's salt is a pale green crystalline solid which is soluble in water. It contains two cations, one of which is  $\text{Fe}^{2+}$ , and one anion which is  $\text{SO}_4^{2-}$ .

The identity of the second cation was determined by heating Mohr's salt with aqueous sodium hydroxide. A colourless gas was evolved which readily dissolved in water giving an alkaline solution.

A green precipitate was also formed.

What are the identities of the gas and the precipitate?

|   | gas           | precipitate              |
|---|---------------|--------------------------|
| A | $\text{NH}_3$ | $\text{Fe}(\text{OH})_2$ |
| B | $\text{NH}_3$ | $\text{Na}_2\text{SO}_4$ |
| C | $\text{SO}_2$ | $\text{Fe}(\text{OH})_2$ |
| D | $\text{SO}_2$ | $\text{Na}_2\text{SO}_4$ |

24 The amount of each gas in the exhaust gases from an internal combustion engine is determined.

The exhaust gases are then passed through the exhaust catalyst, known as a catalytic converter.

Which statements about the effect of the catalytic converter are correct?

- 1 The amount of nitrogen increases.
- 2 The amount of carbon dioxide decreases.
- 3 The amount of carbon monoxide does **not** change.

A 1, 2 and 3      B 1 only      C 2 only      D 3 only

25 In this question, alkenes and cyclic alkanes should be considered.

How many **structural** isomers of  $C_4H_8$  are there?

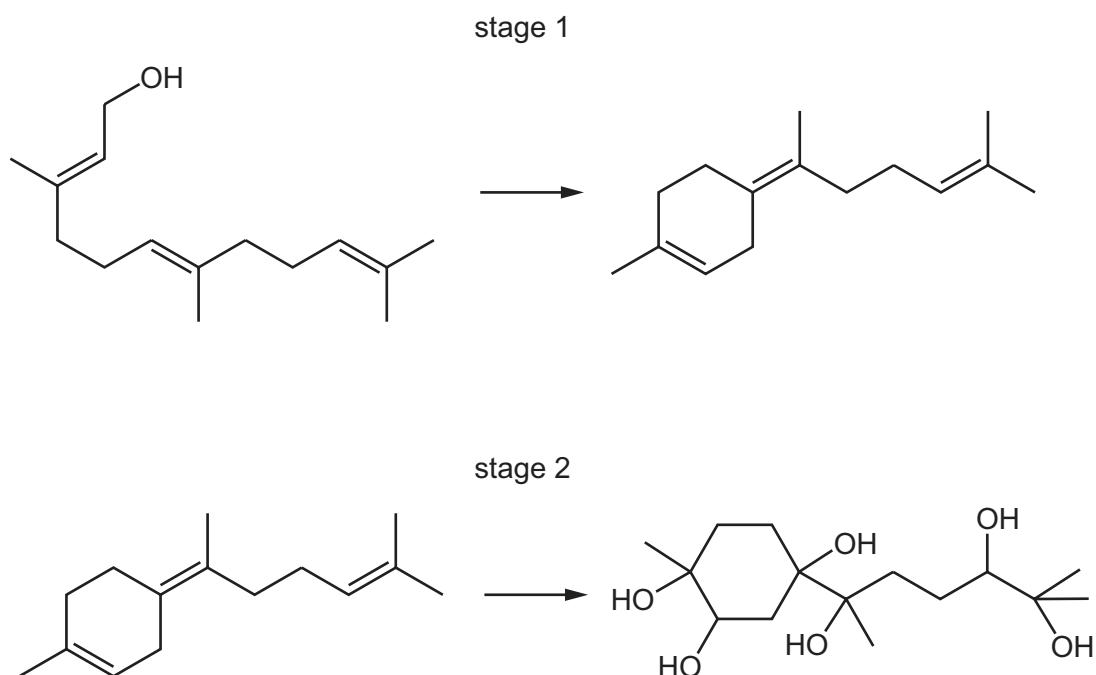
- A 3                      B 4                      C 5                      D 6

26 Which pair of molecules are positional isomers?

- A  $CH_3CH_2CH_2CH_3$  and  $CH_3CH(CH_3)CH_3$   
 B  $CH_3CH_2CHClCH_3$  and  $CH_3CH_2CH_2CH_2Cl$   
 C  $CH_3CH_2CH_2CHO$  and  $CH_3CH_2COCH_3$   
 D  $CH_2CHCH_2CH_3$  and  $CH_3CH_2CHCH_2$

27 The reaction sequence shown takes place in two stages.

The reagent for stage 1 is concentrated  $H_2SO_4$ .



Which type of reaction is carried out in the two stages?

|   | stage 1      | stage 2    |
|---|--------------|------------|
| A | condensation | hydrolysis |
| B | reduction    | hydrolysis |
| C | condensation | oxidation  |
| D | reduction    | oxidation  |

28 Three halogenoalkanes are listed.

- 1 1-bromopropane
- 2 2-chloropropane
- 3 1-bromo-2-chloropropane

Which of these halogenoalkanes could be produced when propene is reacted with a mixture of  $\text{HCl(g)}$  and  $\text{HBr(g)}$ ?

- A** 1, 2 and 3      **B** 1 and 2 only      **C** 1 only      **D** 2 only

29 2-bromomethylpropane is heated with concentrated  $\text{NaOH}$  in ethanol under reflux.

Which statement about the major product of this reaction is correct?

- A** Methylpropene is produced in an elimination reaction.
- B** Methylpropene is produced in a condensation reaction.
- C** Methylpropan-2-ol is produced in an  $\text{S}_{\text{N}}1$  reaction.
- D** Methylpropan-2-ol is produced in an  $\text{S}_{\text{N}}2$  reaction.

30 Which reagents react with ethanol?

- 1  $\text{Na}$
- 2  $\text{SOCl}_2$
- 3  $\text{PCl}_5$

- A** 1, 2 and 3      **B** 1 only      **C** 2 and 3 only      **D** 3 only

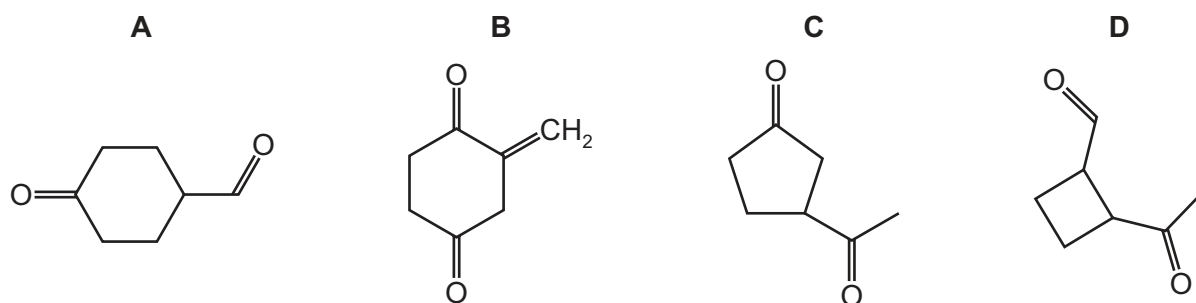
31 Three alkenes are listed.

- 1 hex-1-ene
- 2 cis-hex-2-ene
- 3 trans-hex-3-ene

Which alkenes can be formed by the dehydration of hexan-2-ol?

- A** 1, 2 and 3      **B** 1 and 2 only      **C** 1 only      **D** 2 only

32 Which compound would give positive tests with alkaline  $I_2(aq)$  and with Fehling's reagent?



33 Compound Y is heated with a mild oxidising agent. One of the products of the reaction reacts with hydrogen cyanide, forming 2-hydroxybutanenitrile.

What is compound Y?

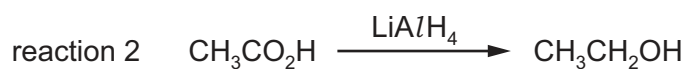
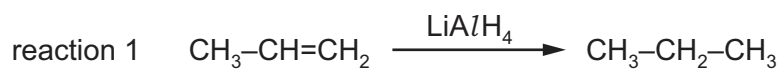
- A** butan-1-ol
- B** butan-2-ol
- C** propan-1-ol
- D** propan-2-ol

34 An ester of structural formula  $CH_3CO_2CH_3$  is heated with aqueous sodium hydroxide.

What are the two organic products of this reaction?

- A** ethanoic acid and methanol
- B** methanoic acid and ethanol
- C** sodium ethanoate and methanol
- D** sodium methanoate and ethanol

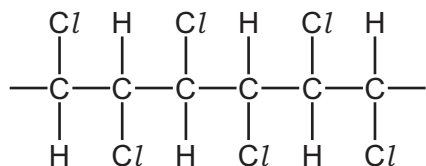
35 A student suggests two uses of  $\text{LiAlH}_4$ .



Which reactions would give the product shown?

- A both reaction 1 and reaction 2
- B reaction 1 only
- C reaction 2 only
- D neither reaction 1 nor reaction 2

36 The diagram shows a section of an addition polymer.



What is the name of this polymer?

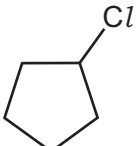
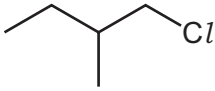
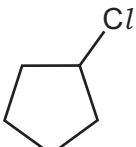
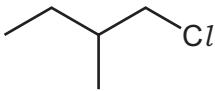
- A poly(1,2-dichloroethane)
- B poly(1,2-dichloroethene)
- C poly(1,2,3-trichloropropane)
- D poly(1,2,3-trichloropropene)

37 0.20 mol of P reacts to form 7.83 g of primary amine Q.

The percentage yield of Q is 45.0%.

[A<sub>r</sub>: C, 12.0; H, 1.0; Cl, 35.5; N, 14.0]

Which row is correct?

|   | reagent and conditions for reaction              | structure of P                                                                      |
|---|--------------------------------------------------|-------------------------------------------------------------------------------------|
| A | KCN in ethanol with heat                         |    |
| B | KCN in ethanol with heat                         |    |
| C | NH <sub>3</sub> in ethanol heated under pressure |   |
| D | NH <sub>3</sub> in ethanol heated under pressure |  |

38 Bromine reacts with ethene in the dark.

Which description of the organic intermediate of this reaction is correct?

- A It has a negative charge.
- B It is a free radical.
- C It is a nucleophile.
- D It is an electrophile.

- 39 When 1-bromopropane is treated in succession with two reagents, X and Y, it produces propanoic acid.

What are reagents X and Y?

|   | reagent X      | reagent Y                                            |
|---|----------------|------------------------------------------------------|
| A | NaOH(aq)       | $\text{H}^+ / \text{Cr}_2\text{O}_7^{2-}(\text{aq})$ |
| B | NaOH(aq)       | $\text{CO}_2$                                        |
| C | KCN in ethanol | $\text{HCl}(\text{aq})$                              |
| D | KCN in ethanol | $\text{NaOH}(\text{aq})$                             |

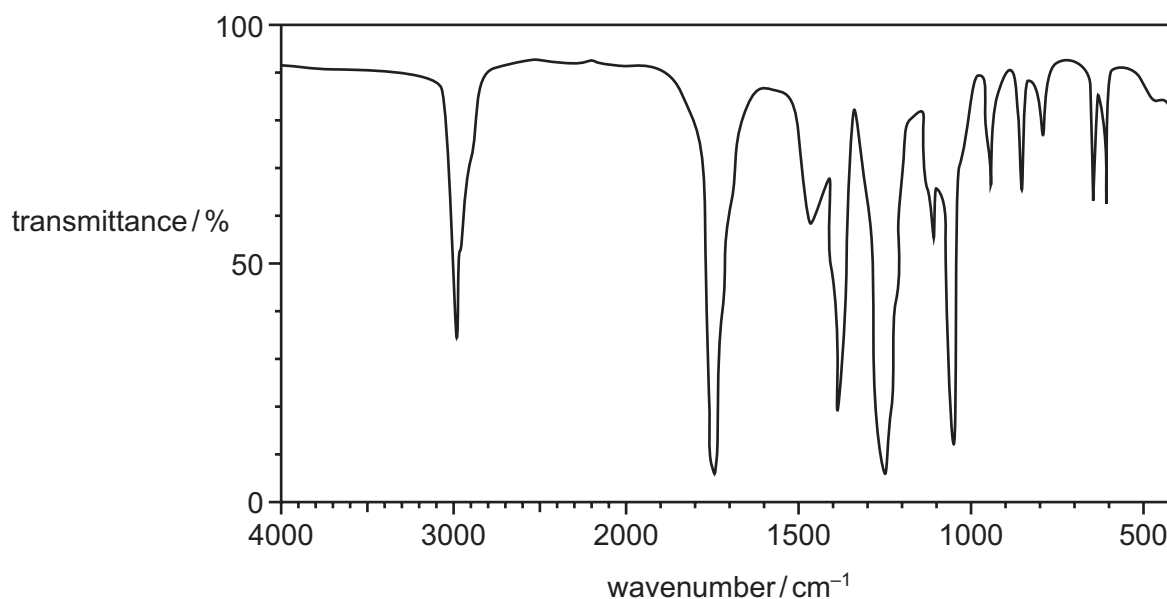
40

| bond | functional groups containing the bond | characteristic infrared absorption range<br>(in wavenumbers)/ $\text{cm}^{-1}$ |
|------|---------------------------------------|--------------------------------------------------------------------------------|
| C–O  | hydroxy, ester                        | 1040–1300                                                                      |
| C=C  | aromatic compound, alkene             | 1500–1680                                                                      |
| C=O  | amide<br>carbonyl, carboxyl<br>ester  | 1640–1690<br>1670–1740<br>1710–1750                                            |
| C≡N  | nitrile                               | 2200–2250                                                                      |
| C–H  | alkane                                | 2850–2950                                                                      |
| N–H  | amine, amide                          | 3300–3500                                                                      |
| O–H  | carboxyl<br>hydroxy                   | 2500–3000<br>3200–3650                                                         |

An organic compound, X, has the composition shown.

C, 54.54%; H, 9.09%; O, 36.37%

The infrared spectrum of X is shown.



What is compound X?

- A** butanoic acid
- B** ethanoic acid
- C** ethyl ethanoate
- D** methyl methanoate

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**Important values, constants and standards**

|                                 |                                                                                                                                           |
|---------------------------------|-------------------------------------------------------------------------------------------------------------------------------------------|
| molar gas constant              | $R = 8.31 \text{ J K}^{-1} \text{ mol}^{-1}$                                                                                              |
| Faraday constant                | $F = 9.65 \times 10^4 \text{ C mol}^{-1}$                                                                                                 |
| Avogadro constant               | $L = 6.022 \times 10^{23} \text{ mol}^{-1}$                                                                                               |
| electronic charge               | $e = -1.60 \times 10^{-19} \text{ C}$                                                                                                     |
| molar volume of gas             | $V_m = 22.4 \text{ dm}^3 \text{ mol}^{-1}$ at s.t.p. (101 kPa and 273 K)<br>$V_m = 24.0 \text{ dm}^3 \text{ mol}^{-1}$ at room conditions |
| ionic product of water          | $K_w = 1.00 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$ (at 298 K (25 °C))                                                             |
| specific heat capacity of water | $c = 4.18 \text{ kJ kg}^{-1} \text{ K}^{-1}$ ( $4.18 \text{ J g}^{-1} \text{ K}^{-1}$ )                                                   |

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The Periodic Table of Elements

| Group                        |                               |                                                                |                               |                               |                                |                               |                                |                              |                                |                               |                               |                               |                               |                               |                                |                              |                             |  |
|------------------------------|-------------------------------|----------------------------------------------------------------|-------------------------------|-------------------------------|--------------------------------|-------------------------------|--------------------------------|------------------------------|--------------------------------|-------------------------------|-------------------------------|-------------------------------|-------------------------------|-------------------------------|--------------------------------|------------------------------|-----------------------------|--|
| 1                            | 2                             | Key                                                            |                               |                               |                                |                               |                                |                              |                                |                               |                               | 13                            | 14                            | 15                            | 16                             | 17                           | 18                          |  |
|                              |                               | atomic number<br>atomic symbol<br>name<br>relative atomic mass |                               |                               |                                |                               |                                |                              |                                |                               |                               | 1<br>H<br>hydrogen<br>1.0     |                               |                               |                                |                              |                             |  |
| 3<br>Li<br>lithium<br>6.9    | 4<br>Be<br>beryllium<br>9.0   |                                                                |                               |                               |                                |                               |                                |                              |                                |                               |                               |                               |                               |                               |                                |                              |                             |  |
| 11<br>Na<br>sodium<br>23.0   | 12<br>Mg<br>magnesium<br>24.3 | 3                                                              | 4                             | 5                             | 6                              | 7                             | 8                              | 9                            | 10                             | 11                            | 12                            |                               |                               |                               |                                |                              |                             |  |
|                              |                               | 21<br>Sc<br>scandium<br>45.0                                   | 22<br>Ti<br>titanium<br>47.9  | 23<br>V<br>vanadium<br>50.9   | 24<br>Cr<br>chromium<br>52.0   | 25<br>Mn<br>manganese<br>54.9 | 26<br>Fe<br>iron<br>55.8       | 27<br>Co<br>cobalt<br>58.9   | 28<br>Ni<br>nickel<br>58.7     | 29<br>Cu<br>copper<br>63.5    | 30<br>Zn<br>zinc<br>65.4      | 31<br>Ga<br>gallium<br>69.7   | 32<br>Ge<br>germanium<br>72.6 | 33<br>As<br>arsenic<br>74.9   | 34<br>Se<br>selenium<br>79.0   | 35<br>Br<br>bromine<br>79.9  | 36<br>Kr<br>krypton<br>83.8 |  |
| 37<br>Rb<br>rubidium<br>85.5 | 38<br>Sr<br>strontium<br>87.6 | 39<br>Y<br>yttrium<br>88.9                                     | 40<br>Zr<br>zirconium<br>91.2 | 41<br>Nb<br>niobium<br>92.9   | 42<br>Mo<br>molybdenum<br>95.9 | 43<br>Tc<br>technetium<br>—   | 44<br>Ru<br>ruthenium<br>101.1 | 45<br>Rh<br>rhodium<br>102.9 | 46<br>Pd<br>palladium<br>106.4 | 47<br>Ag<br>silver<br>107.9   | 48<br>Cd<br>cadmium<br>112.4  | 49<br>In<br>indium<br>114.8   | 50<br>Sn<br>tin<br>118.7      | 51<br>Sb<br>antimony<br>121.8 | 52<br>Te<br>tellurium<br>127.6 | 53<br>I<br>iodine<br>126.9   | 54<br>Xe<br>xenon<br>131.3  |  |
| 55<br>Cs<br>caesium<br>132.9 | 56<br>Ba<br>barium<br>137.3   | 57–71<br>lanthanoids                                           |                               | 73<br>Ta<br>tantalum<br>180.9 | 74<br>W<br>tungsten<br>183.8   | 75<br>Re<br>rhenium<br>186.2  | 76<br>Os<br>osmium<br>190.2    | 77<br>Ir<br>iridium<br>192.2 | 78<br>Pt<br>platinum<br>195.1  | 79<br>Au<br>gold<br>197.0     | 80<br>Hg<br>mercury<br>200.6  | 81<br>Tl<br>thallium<br>204.4 | 82<br>Pb<br>lead<br>207.2     | 83<br>Bi<br>bismuth<br>209.0  | 84<br>Po<br>polonium<br>—      | 85<br>At<br>astatine<br>—    | 86<br>Rn<br>radon<br>—      |  |
| 87<br>Fr<br>francium<br>—    | 88<br>Ra<br>radium<br>—       | 89–103<br>actinoids                                            |                               | 105<br>Db<br>dubnium<br>—     | 106<br>Sg<br>seaborgium<br>—   | 107<br>Bh<br>bohrium<br>—     | 108<br>Hs<br>hassium<br>—      | 109<br>Mt<br>meitnerium<br>— | 110<br>Ds<br>darmstadtium<br>— | 111<br>Rg<br>roentgenium<br>— | 112<br>Cn<br>copernicium<br>— | 113<br>Nh<br>nihonium<br>—    | 114<br>Fl<br>flerovium<br>—   | 115<br>Mc<br>moscovium<br>—   | 116<br>Lv<br>livermorium<br>—  | 117<br>Ts<br>tennessine<br>— | 118<br>Og<br>oganeson<br>—  |  |